

Bond Polarity, it's an "electron tug of war"

Look up the electronegativities of the elements involved in the following compounds and write them under each element. Indicate the type of bond that holds them together by writing **P** (polar), **NP** (nonpolar) or **I** (Ionic). If they exist, label the atoms with partial charges (**δ^-** and **δ^+**). Finally, calculate the electronegativity (**END**) difference for each compound.

Compound	Type of Bonds	Electronegativity Difference (END)
oxygen difluoride OF₂		
nitrogen trichloride NCl₃		
hydrogen bromide HBr		
carbon tetrafluoride CF₄		
carbon disulfide CS₂		
nitrogen N₂		
calcium iodide CaI₂		

2) Based on electronegativity difference, which of the molecular compounds above has the greatest ionic character? Which has the least ionic character?