

Don't be square. Be tetrahedral, man!

Draw the electron dot diagrams of the following molecules. **Then, based on the # bonding pairs and lone pairs around the central atom, draw the 3D structures** of the following molecules.

<p>1) SeBr₂ selenium dibromide</p>	<p>4) SiF₄ silicon tetrafluoride</p>
<p>2) NI₃ nitrogen triiodide</p>	<p>5) GeS₂ germanium disulfide</p>
<p>3) HBr hydrogen bromide</p>	<p>6) F₂ fluorine</p>