From last nite's HW,

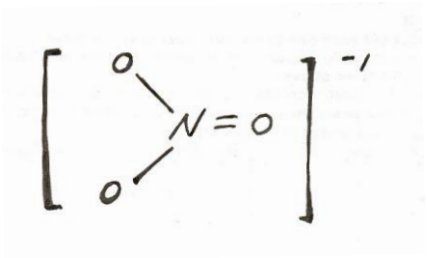
(e) lead(II) nitrate Pb_3N_2 ????

No sir! It's ate not ide.

GO TO TABLE E**: Selected Polyatomic Ions**

Aim: How do we determine the formulas of compounds containing "polys"?

polyatomic ion- many atoms with a charge

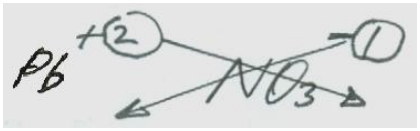


"**nitrate**" ion

I. CRISS-CROSS with "POLYS"

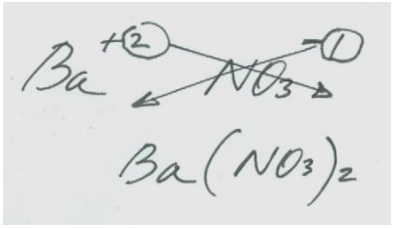
1) To get the formula, "criss-cross" the oxidation states treating the poly as one atom.

(e) lead(II) nitrate



2) If there is more than one poly, use **parentheses**.

barium nitrate



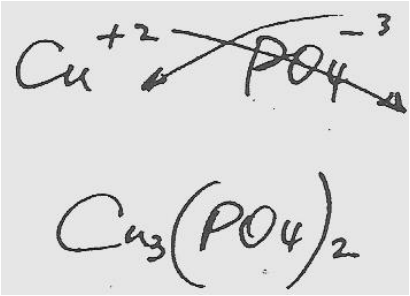
This means that there is **one** Ba^{+2} ion and **two** ClO_3^{-1} ions in a unit of this compound.

"cut & paste"

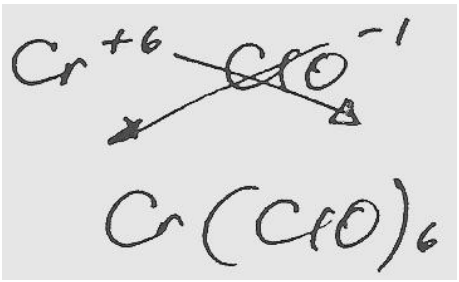
More Practice

given name, write formula:

copper (II) phosphate



chromium (VI) hypochlorite



Let's go to today's handout: Writng Formulas from Names

WRITING FORMULAS FROM NAMES

Write the formulas of the following compounds.

1. ammonium phosphate — $(NH_4)_3PO_4$
2. iron (II) oxide — FeO
3. iron (III) oxide — Fe_2O_3
4. * carbon monoxide —
5. calcium chloride — $CaCl_2$
6. potassium nitrate — KNO_3
7. magnesium hydroxide — $Mg(OH)_2$
8. aluminum sulfate — $Al_2(SO_4)_3$
9. copper (II) sulfate — $CuSO_4$
10. lead (IV) chromate — $Pb(CrO_4)_2$
11. * diphosphorus pentoxide —
12. potassium permangdnate — $KMnO_4$
13. sodium hydrogen carbonate — $NaHCO_3$
14. zinc nitrate — $Zn(NO_3)_2$
15. aluminum sulfite — $Al_2(SO_3)_3$