holy moley 2

1. The sum of the atomic masses of the atoms in one molecule of $C_3H_6Br_2$ is called the

   1) formula mass
   2) isotopic mass
   3) percent abundance
   4) percent composition

2. The gram-formula mass of NO_2 is defined as the mass of

   1) one mole of NO_2
   2) one molecule of NO_2
   3) two moles of NO
   4) two molecules of NO

3. A 1.0-mole sample of krypton gas has a mass of

   1) 19 g
   2) 36 g
   3) 39 g
   4) 84 g

4. What is the gram-formula mass of $Ca_3(PO_4)_2$?

   1) 248 g/mol
   2) 263 g/mol
   3) 279 g/mol
   4) 310. g/mol

5. The molar mass of $Ba(OH)_2$ is

   1) 154.3 g
   2) 155.3 g
   3) 171.3 g
   4) 308.6 g

6. What is the gram molecular mass of 1 mole of $C_3H_5(OH)_3$?

   1) 48 g
   2) 58 g
   3) 74 g
   4) 92 g

7. Which sample contains a mole of atoms?

   1) 23 g Na
   2) 24 g C
   3) 42 g Kr
   4) 78 g K

8. One mole of O_2 has approximately the same mass as one mole of

   1) CH_4
   2) S
   3) LiH
   4) Cl_2

holy moley 2

9. The total number of moles represented by 20 grams of $CaCO_3$ is

1) 1 2) 2 3) 0.1 4) 0.2

10. What is the mass in grams of 2.0 moles of NO_2?

1) 92 2) 60. 3) 46 4) 30.

11. Which sample contains the same number of atoms as 24 grams of carbon?

1) 80. g Ar 2) 24 g Mg 3) 10. g Ne 4) 4.0 g He

12. What is the total mass of oxygen in 1.00 mole of $Al_2(CrO_4)_3$?

1) 192 g 2) 112 g 3) 64.0 g 4) 48.0 g

13. A sample of a compound contains 65.4 grams of zinc, 12.0 grams of carbon, and 48.0 grams of oxygen. What is the mole ratio of zinc to carbon to oxygen in this compound?

1) 1:1:2 2) 1:1:3 3) 1:4:6 4) 5:1:4

14. The gram molecular mass of CO_2 is the same as the gram molecular mass of

1) CO 2) SO_2 3) C_2H_6 4) C_3H_8

15. The number of moles of molecules in a 12.0-gram sample of Cl_2 is

1) $\frac{12.0}{35.5}$ mole 2) $\frac{12.0}{71.0}$ mole 3) 12.0 moles 4) $\frac{12.0 \times 35.5}{\text{moles}}$