6) Sulfuric acid, $H_2SO_4(aq)$, can be used to neutralize barium hydroxide, $Ba(OH)_2(aq)$. What is the formula for the salt produced by this neutralization?

A) $BaSO_3$ B) $BaSO_2$ C) BaS D) $BaSO_4$

7) Which reactants form the salt $CaSO_4(s)$ in a neutralization reaction?

A) $H_2SO_4(aq)$ and $Ca(OH)_2(aq)$
B) $H_2SO_3(aq)$ and $Ca(NO_3)_2(aq)$
C) $SO_2(g)$ and $CaO(s)$
D) $H_2S(g)$ and $Ca(ClO_4)_2(s)$

Refer to Tables K, L, E & Periodic Table to complete the following equations.

1) Nitrous Acid + Lithium Hydroxide

a) Molecular equation:

b) Ionic equation:

c) Net ionic equation:

d) Name of salt formed:

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2) Sulfurous Acid + Barium Hydroxide

a) Molecular equation:

b) Ionic equation:

c) Net ionic equation:

d) Name of salt formed:

SHOW ALL WORK: Circle & label keywords, show correct numerical set up & calculate

$$M_A \cdot V_A = M_B \cdot V_B$$

1) How many ml of 1.5 M NaOH are needed to completely neutralize 75 ml of 3.0 M HCl?

2) How many ml of 0.25 M HNO_3 are needed to completely neutralize 55 ml of 0.50 M KOH?

MORE CHALLENGING QUESTIONS

1) How many ml of 1.0 M NaOH are needed to completely neutralize 50. ml of 1.0 M H_2SO_4? (**Note**: 2 H in H_2SO_4)

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2) Phosphoric Acid + Potassium Hydroxide

a) Molecular equation:

b) Ionic equation:

c) Net ionic equation:

d) Name of salt formed: